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## ADMISSIBLE EQUATIONS OF STATE FOR IMMISCIBLE AND MISCIBLE MIXTURES <sup>\*,\*\*</sup>

GLORIA FACCANONI<sup>1</sup> AND HÉLÈNE MATHIS<sup>2</sup>

**Abstract.** This paper addresses the construction of admissible Equations of State (EoS) for compressible two-phase flows. We investigate two approaches. In the first one, the mixture is treated as a single fluid with a complex thermodynamic. Most of the time the available EoS are determined experimentally and are often incomplete EoS, *i.e.* we know only the pressure as a function of the volume and the temperature. We present here a general framework to compute a complete EoS based on such an incomplete EoS. In the second approach, each phase is depicted by its own EoS. Following the Gibbs formalism, the mixture entropy is the sum of the phasic entropies which achieves its maximum at equilibrium. Depending on the miscibility of the mixture, one gets different geometrical properties on the resulting mixture entropy. Eventually we address the coupling of mixture EoS with the dynamic of the fluid. Homogeneous Equilibrium and Relaxation Models (HEM and HRM) are introduced for an immiscible and a miscible two-phase mixture. Hyperbolicity is ensured taking advantage of the concavity properties of the mixture entropies.

**Key-words.** Two-phase flows, equation of state, thermodynamic of equilibrium, phase transition, homogeneous equilibrium and relaxation model, hyperbolicity.

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### 1. INTRODUCTION

Compressible multiphase flows are involved in many industrial applications, notably in nuclear framework, for instance in loss of flow or loss of coolant accidents in pressurized water reactor. For the purpose of CFD calculations, thermodynamic properties must be calculated over a wide range of states. Thus it is mandatory to provide a consistent Equation of State (EoS) to give a qualitatively and quantitatively correct material description. This EoS should represent phase equilibria over wide temperature and pressure ranges, in pure region (liquid or vapor) as well as in the VLE (Vapor-Liquid Equilibrium) region.

The construction of coherent EoS raises several difficult questions:

- (1) how to obtain an accurate (and useful in CFD) description of a pure region?

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- (2) how to compute a given thermodynamic state by performing calculations with different couples of input variables?
- (3) how to describe accurately the transition from a pure region to the VLE region?
- (4) CFD requires extremely fast algorithms to compute thermodynamic properties because such algorithms are frequently used in the inner iteration cycles of the process calculations. How to balance accuracy and reasonable computing times?

A classic strategy for the last question (4) is to use an analytical EoS. So the question is to construct such an analytical EoS while respecting constraints 1-3.

Focusing on a macroscopic description of the compressible two-phase flow, we consider that the dynamic of the fluid is depicted by an Euler-like system. The two-phase mixture is treated as a single fluid with a complex thermodynamic.<sup>1</sup>

In order to define the complex EoS, we consider two main approaches.

- The first possibility is to use an analytical EoS, valid in a wide range of temperatures and pressures. Such equations are often deduced from measurements and are usually incomplete in the sense that they define solely the pressure  $p$  as a function of the volume  $\tau$  and the temperature  $T$ . However, in the CFD calculations (based on the Euler equations), it is necessary to specify the pressure as a function of the volume and the specific internal energy  $e$ . To obtain these relation, an idea is to construct an entropy function  $(\tau, e) \mapsto s$ , which is actually a complete EoS. It is then possible to derive all incomplete EoS, including the pressure and the temperature. We present this general construction in Section 2. The section begins by the definitions of the different thermodynamic potentials both in the extensive and intensive settings. Then in paragraph 2.2 we build a complete EoS starting from the incomplete law  $(\tau, T) \mapsto p$ . The principle is to determine a temperature  $(\tau, e) \mapsto T$  and thus an entropy  $(\tau, e) \mapsto s$  thermodynamically consistent, *i.e.* they satisfy the first principle of the Thermodynamics and a differential relation of the form  $Tds = de + pd\tau$ .
- The second approach we present consists in assuming that each phase is depicted by its own complete EoS (for instance its entropy). Out of equilibrium the mixture entropy is the sum of the phasic entropies. According to the second principle of Thermodynamics, the mixture entropy reaches its maximum at equilibrium, under the constraints of the mass and energy conservation and a volumic constraint. Depending on the miscibility of the mixture, the latter constraint changes leading to different geometrical properties of the mixture entropy. We present in Section 3 the construction principle in the case of binary mixture. Note that it can be extend to multiphase mixture, see for instance [21, 29] for three-phase framework. A particular attention is paid on the characterization of the thermodynamic equilibrium and on the concavity properties of the mixture entropy.

In Section 4, we address the coupling of the admissible EoS computed in Section 3 with the dynamic of the fluid. We present two Homogeneous Equilibrium Models (HEM) depending on the miscibility of the two-phase mixture. They consist in Euler type models with the closure laws computed in Section 3. The concavity properties of the mixture entropies ensure that the models are hyperbolic. However the closure laws is often difficult to compute and may not be analytical, even for simple phasic EoS. It is then convenient to approximate the model by a relaxation procedure, see Section 5. One obtains an Homogeneous Relaxation Model (HRM), with extra equations and relaxation source terms toward the thermodynamic equilibrium, which comply with the entropy growth criterion. Again we present two models according to the miscibility of the two-phase mixture.

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<sup>1</sup>Note that a way to avoid this definition of the complex EoS is to consider a two-fluid description, which treats each phase as a separated fluid depicted by balance laws of phasic mass conservation, phasic momentum and energy which are coupled through exchanging terms and nonconservative terms. Then the modelling difficulties are shifted on the definition of the exchanging terms, see [2, 17].

## 2. SINGLE FLUID THERMODYNAMICS

### 2.1. Extensive and intensive notations

Let consider a fluid of mass  $M \geq 0$  and internal energy  $E \geq 0$  occupying a volume  $V \geq 0$ . Let us denote  $\widetilde{W} \stackrel{\text{def}}{=} (M, V)$  and  $W \stackrel{\text{def}}{=} (V, M, E)$ . As the fluid is homogeneous and at rest, its thermodynamic behaviour is entirely described by its entropy function:

$$\begin{aligned} S: (\mathbb{R}^+)^3 &\rightarrow \mathbb{R} \\ W &\mapsto S(W). \end{aligned}$$

This entropy function  $S$  is uniquely and smoothly determined by the state vector  $W \in (\mathbb{R}^+)^3$ . Standard Thermodynamic requires that  $S$  is a concave function of  $W$ . Then it is classic to extend it by  $-\infty$  outside the close convex cone  $(\mathbb{R}^+)^3$

$$S(W) = \begin{cases} S(W), & W \in (\mathbb{R}^+)^3, \\ -\infty, & \text{elsewhere.} \end{cases}$$

We adopt the assumptions stated in [4] and [10]:

**Assumption 1.** (Extensive assumptions) The entropy  $S: (\mathbb{R}^+)^3 \rightarrow \mathbb{R} \cup \{-\infty\}$  is such that

- (i) the set of admissible states  $C \stackrel{\text{def}}{=} \{W \in (\mathbb{R}^+)^3 \mid S(W) > -\infty\}$  is a non-empty close convex domain,
- (ii)  $S$  is a concave function of  $W \in C$ ,
- (iii)  $S$  is *extensive*, *i.e.* Positively Homogeneous of degree 1 (PH1), that is

$$\forall \lambda \in \mathbb{R}_*^+, \forall W \in C, \quad S(\lambda W) = \lambda S(W),$$

- (iv)  $S$  is upper semi-continuous (u.s.c.) that is

$$\forall W_0 \in C, \quad \lim_{W \rightarrow W_0} \sup S(W) \leq S(W_0),$$

- (v)  $S$  is of class  $\mathcal{C}^2$  on  $C$  and its partial derivative with respect to the internal energy  $E$  is strictly positive

$$\forall W \in C, \quad \frac{\partial S}{\partial E} > 0.$$

Assumptions (ii) and (iii) are equivalent to assume  $(-S)$  sub-linear [31] or equivalently that the epigraph of  $(-S)$  is a closed set in  $\mathbb{R}^3 \times \mathbb{R}$ . The set  $C$  is not necessarily convex, as it has been studied in [14, 23]. Observe that the extensive entropy  $S$  cannot be strictly concave since it is PH1.

The fundamental extensive thermodynamic relation is

$$TdS = dE + pdV - \mu dM, \tag{1}$$

where  $W \mapsto T$  is the temperature,  $W \mapsto p$  the pressure and  $W \mapsto \mu$  the chemical potential given by

$$T = \left( \frac{\partial S}{\partial E} \Big|_{M,V} \right)^{-1}, \quad p = T \frac{\partial S}{\partial V} \Big|_{M,E}, \quad \mu = -T \frac{\partial S}{\partial M} \Big|_{V,E}. \tag{2}$$

As derivatives of a PH1 function, this functions are PH0 or intensive potentials. Since  $S$  is PH1, it satisfies the Euler's relation  $\nabla_W S \cdot W^\top = S$  which leads to another characterization of the chemical potential

$$M\mu = -TS + pV + E. \tag{3}$$

**Remark 2.1.** Following [9], the formula  $S = S(W)$  is called the *fundamental equation* (or *complete equation of state*) of our system, and by definition contains all thermodynamic informations. An identity involving other derived quantities (such  $T$  or  $p$  for instance) is an *incomplete equation of state*, which typically does not contain all thermodynamic information.

### 2.1.1. Intensive properties — specific description

By homogeneity of the extensive entropy function, one may introduce the *specific entropy*  $s$  as

$$s \stackrel{\text{def}}{=} \frac{1}{M} S(W) \stackrel{\text{PH1}}{=} S\left(\frac{W}{M}\right) = S\left(\frac{V}{M}, 1, \frac{E}{M}\right). \quad (4)$$

It corresponds to a *complete* EoS which is uniquely determined by the *specific volume*  $\tau \stackrel{\text{def}}{=} V/M$  and the *specific internal energy*  $e \stackrel{\text{def}}{=} E/M$ :

$$\begin{aligned} s: (\mathbb{R}^+)^2 &\rightarrow \mathbb{R} \\ (\tau, e) &\mapsto s(\tau, e). \end{aligned} \quad (5)$$

The *fundamental intensive thermodynamic relation* is

$$T ds = de + p d\tau, \quad (6)$$

where the temperature  $T$  and the pressure  $p$  are now functions of  $w \stackrel{\text{def}}{=} (\tau, e)$ :

$$w \mapsto T = \left( \frac{\partial s}{\partial e} \Big|_{\tau} \right)^{-1}, \quad w \mapsto p = T \frac{\partial s}{\partial \tau} \Big|_e \quad (7)$$

and the chemical potential satisfy

$$\mu = -Ts + p\tau + e.$$

Assumptions 1 admit their intensive counterpart.

**Assumption 2.** (Intensive specific assumptions) The specific entropy  $s$  is such that

- the set of admissible states  $c \stackrel{\text{def}}{=} \{ w \in (\mathbb{R}^+)^2 \mid s(w) > -\infty \}$  is a non-empty close convex domain,
- $s$  is of class  $\mathcal{C}^2$  on  $c$  such that

$$T = \left( \frac{\partial s}{\partial e} \Big|_{\tau} \right)^{-1} > 0, \quad (8)$$

- its Hessian matrix is negative definite, that is

$$\det(H_s) \stackrel{\text{def}}{=} \frac{\partial^2 s}{\partial \tau^2} \Big|_e \frac{\partial^2 s}{\partial e^2} \Big|_{\tau} - \left( \frac{\partial^2 s}{\partial \tau \partial e} \right)^2 > 0, \quad \frac{\partial^2 s}{\partial \tau^2} \Big|_e < 0, \quad \frac{\partial^2 s}{\partial e^2} \Big|_{\tau} < 0, \quad (9)$$

which implies that  $s$  is strictly concave (but the converse is wrong).<sup>2</sup>

### 2.1.2. Intensive properties — volumic description

The homogeneity of the extensive entropy allows to introduce also the *volumic entropy*  $\sigma$  satisfying

$$\sigma \stackrel{\text{def}}{=} \frac{1}{V} S(W) \stackrel{\text{PH1}}{=} S\left(\frac{W}{V}\right) = S\left(1, \frac{M}{V}, \frac{E}{V}\right). \quad (10)$$

<sup>2</sup>Note that the concavity of  $s$  may be deduced from the concavity of the extensive entropy  $S$  and its homogeneity.

It is also a *complete EoS* uniquely determined by the density  $\rho \stackrel{\text{def}}{=} M/V = 1/\tau > 0$  and the *volumic* internal energy  $\varepsilon \stackrel{\text{def}}{=} E/V = \rho e$  such that

$$\begin{aligned} \sigma: (\mathbb{R}^+)^2 &\rightarrow \mathbb{R} \\ (\rho, \varepsilon) &\mapsto \sigma(\rho, \varepsilon). \end{aligned} \quad (11)$$

As a function of  $\tilde{w} \stackrel{\text{def}}{=} (\rho, \varepsilon)$ , it complies with the following total differential

$$Td\sigma = d\varepsilon - \mu d\rho, \quad (12)$$

where the temperature  $T$  and the chemical potential  $\mu$  are now functions of  $\tilde{w}$ :

$$\tilde{w} \mapsto T = \left( \frac{\partial \sigma}{\partial \varepsilon} \Big|_{\rho} \right)^{-1}, \quad \tilde{w} \mapsto \mu = -T \frac{\partial \sigma}{\partial \rho} \Big|_{\varepsilon}, \quad (13)$$

and the pressure satisfy

$$p = T\sigma + \rho\mu - \varepsilon. \quad (14)$$

**Assumption 3.** (Intensive volumic assumptions) The volumic entropy  $\sigma$  is such that

- the set of admissible states  $\tilde{c} = \{\tilde{w} \in (\mathbb{R}^+)^2 | \sigma(\tilde{w}) > -\infty\}$  is a non-empty close convex domain,
- $\sigma$  is of class  $\mathcal{C}^2$  on  $\tilde{c}$  such that

$$T = \left( \frac{\partial \sigma}{\partial \varepsilon} \Big|_{\rho} \right)^{-1} > 0, \quad (15)$$

- its Hessian matrix is negative definite, that is

$$\det(H_{\sigma}) \stackrel{\text{def}}{=} \frac{\partial^2 \sigma}{\partial \rho^2} \Big|_{\varepsilon} \frac{\partial^2 \sigma}{\partial \varepsilon^2} \Big|_{\rho} - \left( \frac{\partial^2 \sigma}{\partial \rho \partial \varepsilon} \right)^2 > 0, \quad \frac{\partial^2 \sigma}{\partial \rho^2} \Big|_{\varepsilon} < 0, \quad \frac{\partial^2 \sigma}{\partial \varepsilon^2} \Big|_{\rho} < 0, \quad (16)$$

which implies that  $\sigma$  is strictly concave (but the converse is wrong).<sup>3</sup>

## 2.2. From an incomplete to a complete EoS

It is very common to specify the EoS of a material by providing the pressure as a function of the volume and the temperature:  $(\tau, T) \mapsto p$ . As noted in Remark 2.1, such an expression does not embody the full range of thermodynamic informations available from the fundamental equation  $S = S(W)$  (or  $s = s(w)$ ). However, in the CFD calculations (based on the Euler equations), one need to specify the pressure as a function of the volume and the internal energy:  $(\tau, e) \mapsto p$ . To do so, an idea is to construct a complete EoS in the form of a specific entropy  $(\tau, e) \mapsto s$ , based on the knowledge of the incomplete EoS  $(\tau, T) \mapsto p$ . Once the complete EoS is built, all other quantities can be computed thanks to the fundamental thermodynamic relation (6) and the Maxwell relations (see appendix A).

### 2.2.1. General construction

In this section, we focus on the specific quantities and present a general method to deduce a complete EoS  $(\tau, e) \mapsto s$  from an incomplete EoS  $(\tau, T) \mapsto p$ . This construction involves determining a temperature  $(\tau, e) \mapsto T$  and thus a specific entropy  $(\tau, e) \mapsto s$ , which are thermodynamically consistent in the sense that it complies with the fundamental relation (6).

Let us consider a general incomplete EoS  $(\tau, T) \mapsto p$  of the form

$$(\tau, T) \mapsto p = \frac{rT}{\tau - b} - a\mathcal{Q}(\tau), \quad (17)$$

<sup>3</sup>Note that the concavity of  $\sigma$  may be deduced from the concavity of the extensive entropy  $S$  by homogeneity, or otherwise, from the concavity of the specific entropy  $s$ , since  $\sigma = \rho s$ .

where  $a$  and  $b \geq 0$  are material-specific constants and  $r > 0$  is the universal gas constant. The choice  $\mathcal{Q}(\tau) = 1$  corresponds to the (Noble-Abel)-Stiffened-Gas (NASG) law [5, 26, 27], while the choice  $\mathcal{Q}(\tau) = \tau^{-2}$  corresponds to the van der Waals law [4, 25].

Since the specific entropy  $s$  is a total derivative which satisfies the fundamental relation (6), its cross derivatives are equal:

$$\left. \frac{\partial \left( \frac{p}{T} \right)}{\partial e} \right|_{\tau} = \left. \frac{\partial \left( \frac{1}{T} \right)}{\partial \tau} \right|_e. \quad (18)$$

Using the pressure definition (17), one can exhibit an hyperbolic PDE satisfied by the temperature  $(\tau, e) \mapsto T$ :

$$\partial_{\tau} T + a \mathcal{Q}(\tau) \partial_e T = 0. \quad (19)$$

Let introduce the characteristic curves  $\tau \mapsto \chi(\tau)$  associated with the transport equation (19). In the  $(\tau, e)$ -plane they are solutions of the Cauchy problem

$$\begin{cases} \chi'(\tau) = a \mathcal{Q}(\tau), \\ \chi(\tau_0) = \chi_0, \end{cases} \quad (20)$$

that is

$$\chi(\tau) = \chi_0 + a \int_{\tau_0}^{\tau} \mathcal{Q}(x) dx. \quad (21)$$

The temperature  $\mathcal{T} \stackrel{\text{def}}{=} T(\tau, e = \chi(\tau))$  is constant along the characteristic curves, hence it holds  $\mathcal{T}(\tau) = \mathcal{T}(\tau_0)$  and yields

$$T(\tau, e) = T \left( \tau_0, e - a \int_{\tau_0}^{\tau} \mathcal{Q}(x) dx \right). \quad (22)$$

For illustrative purposes, we choose the temperature as

$$T(\tau, e) = \frac{e - a \int_{\tau_0}^{\tau} \mathcal{Q}(\tau) d\tau}{c_v}, \quad (23)$$

with  $c_v$  the specific heat (which is assumed constant here for sake of simplicity). Finally a complete EoS can be obtained by integrating the fundamental relation (6)

$$ds = \frac{p}{T} d\tau + \frac{1}{T} de = \left( \frac{c_v}{e - a \int \mathcal{Q}(\tau) d\tau} p(\tau, e) \right) d\tau + \frac{c_v}{e - a \int \mathcal{Q}(\tau) d\tau} de. \quad (24)$$

Using (8) and treating  $\tau$  as a constant, by integrating the relation (24) with respect to  $e$  gives

$$s(\tau, e) = \int \frac{1}{T} de + \omega(\tau) = \int \frac{c_v}{e - a \int \mathcal{Q}(\tau) d\tau} de + \omega(\tau) = c_v \ln \left( e - a \int \mathcal{Q}(\tau) d\tau \right) + \omega(\tau). \quad (25)$$

In order to determine the function  $\omega$ , one has to derive (25) with respect to  $\tau$ . It holds

$$\left. \frac{\partial s}{\partial \tau} \right|_e = \frac{p}{T} = \frac{r}{\tau - b} - \frac{c_v}{e - a \int \mathcal{Q}(\tau) d\tau} a \mathcal{Q}(\tau).$$

Using the general form of the pressure (17) and the temperature definition (23), by identification, the function  $\omega(\tau)$  satisfies  $\omega'(\tau) = \frac{r}{\tau - b}$  and consequently  $\omega(\tau) = r \ln(\tau - b)$ . Finally the specific entropy reads

$$s(\tau, e) = c_v \ln \left( e - a \int \mathcal{Q}(\tau) d\tau \right) + r \ln(\tau - b) + s_0, \quad (26)$$

where  $s_0$  is a numerical constant of integration.

We underline that, given  $\tau \mapsto \mathcal{Q}(\tau)$ , the definition of the temperature  $T$  is not unique (and, as a consequence, the specific entropy is not unique as well).

### 2.2.2. Applications

Considering the aforementioned choices of  $\mathcal{Q}(\tau)$ , one obtains the following complete EoS:

- the Noble-Abel-Stiffened-Gas law corresponds to  $\mathcal{Q}(\tau) = 1$  which leads to

$$s(\tau, e) = c_v \ln(e + a\tau + \text{const.}) + r \ln(\tau - b) + s_0.$$

Getting  $\text{const.} = -q + ab$  where  $q$  is a constant, we recover the Noble-Abel-Stiffened-Gas EoS described in [5, 26, 27]. This specific entropy  $s$  is a strictly concave function of  $(\tau, e)$  and complies with Assumptions 2.

- The van der Waals law is obtained setting  $\mathcal{Q}(\tau) = \tau^{-2}$  and gives

$$s(\tau, e) = c_v \ln\left(e + \frac{a}{\tau} + \text{const.}\right) + r \ln(\tau - b) + s_0.$$

Classically  $\text{const.} = 0$ , see [4, 25]. However  $(\tau, e) \mapsto s$  is not concave everywhere. Thus the van der Waals specific entropy fits into the foregoing framework (Assumption 2) only if we restrict attention to regions where  $(\tau, e) \mapsto s$  is concave. If we apply the well-know *Maxwell area construction*, which is equivalent to replace  $s$  by its concave hull (the smallest concave function greater than or equal to  $s$  in some region), the graph contains a line and this geometrical property is related to the physical interpretation of phase transition (see [4]). In [15, 28] the authors investigate the link between the Legendre Transform (LT) and the sup-convolution and the concave hull. The sup-convolution is transformed into an addition by the LT and applying twice a LT to any function is an elegant way to compute its concave hull. A fast LT algorithm is proposed in [15] in order to compute tabulated Equation of State of mixture with application to the van der Waals law.

## 3. COMPUTING ADMISSIBLE EOS FOR TWO-PHASE MIXTURE

We consider an ideally isolated system of fixed mass  $M \geq 0$ , volume  $V \geq 0$  and internal energy  $E \geq 0$ , composed of two-phases  $A$  and  $B$  which are entirely described by their entropy functions  $S_k$ , function of their masses  $0 \leq M_k \leq M$ , volumes  $0 \leq V_k \leq V$  and internal energies  $0 \leq E_k \leq E$ ,  $k \in \mathbb{K} \stackrel{\text{def}}{=} \{A, B\}$ . In the following we denote  $W_k \stackrel{\text{def}}{=} (V_k, M_k, E_k)$  and  $\widetilde{W}_k \stackrel{\text{def}}{=} (M_k, E_k)$ .

Mass and energy conservations ensure that

$$M = M_A + M_B, \tag{27}$$

$$E = E_A + E_B. \tag{28}$$

Depending on the miscibility of the mixture, the volume constraint should differ:

- as the two phases are **immiscible**, they are separated by an interface, leading to

$$V = V_A + V_B, \tag{29}$$

so that  $W = W_A + W_B$ ;

- when the two phases are **miscible**, we assume they are completely intimate and no interface arises, meaning that they behave like ideal gases in the sense that

$$V = V_A = V_B, \tag{30}$$

so that  $\widetilde{W} = \widetilde{W}_A + \widetilde{W}_B$ .



**Remark 3.1.** Note that in both cases, we assume that no vacuum appears in the model. Otherwise one should consider that  $V \geq V_A + V_B$ .

**Remark 3.2.** As the two phases are immiscible, one should consider  $V_A + V_B \leq V$ . As explained in [16], if the phasic pressures  $p_k$ ,  $k \in \mathbb{K}$ , are always strictly positive, then the constraint is necessarily saturated and boils down to  $V = V_A + V_B$ . In a perfectly miscible case, one should consider  $V_A \leq V$  and  $V_B \leq V$ . Again as the phasic pressures and the temperatures are non-negative, the constraint is saturated in the sense that  $V = V_A = V_B$ .

### 3.1. Characterization of the thermodynamic equilibrium

Since the fluid system is ideally isolated, the second principle of Thermodynamic states that all internal processes lead to an increase on the total extensive entropy of the system. When the extensive entropy reaches its maximum value, the system is at thermodynamic equilibrium. In this section we address the definition of the extensive equilibrium entropy of the mixture depending on its miscibility.

Out of equilibrium, the mixture entropy  $\mathcal{S}$  is the sum of the phasic entropies  $S_k$ . For  $(W_A, W_B) \in C_A \times C_B$ , it reads

$$\mathcal{S}(W_A, W_B) = S_A(W_A) + S_B(W_B) = S_A(V_A, M_A, E_A) + S_B(V_B, M_B, E_B). \quad (31)$$

**Definition 3.3.** (Extensive **equilibrium** entropy of an **immiscible** mixture) Let  $W \stackrel{\text{def}}{=} (V, M, E) \in (\mathbb{R}^+)^3$  be the state vector of the two-phase **immiscible** mixture. The equilibrium entropy is:

$$S^{\text{immi}}(W) \stackrel{\text{def}}{=} \max_{(W_A, W_B) \in \Omega^{\text{immi}}} \mathcal{S}(W_A, W_B), \quad (32)$$

where

$$\begin{aligned} \Omega^{\text{immi}} &\stackrel{\text{def}}{=} \{ (W_A, W_B) \in C_A \times C_B \mid (27), (28) \text{ and } (29) \text{ hold} \} \\ &= \{ (W_A, W_B) \in C_A \times C_B \mid W = W_A + W_B \}. \end{aligned}$$

**Definition 3.4.** (Extensive **equilibrium** entropy of a **miscible** mixture) Let  $V = V_0 > 0$  fixed. Let  $\widetilde{W} \stackrel{\text{def}}{=} (M, E) \in (\mathbb{R}^+)^2$  be the state vector of the two-phase **miscible** mixture. The equilibrium entropy is

$$S^{\text{mi}}(V_0, \widetilde{W}) \stackrel{\text{def}}{=} \max_{(\widetilde{W}_A, \widetilde{W}_B) \in \Omega^{\text{mi}}} \mathcal{S}((V_0, \widetilde{W}_A), (V_0, \widetilde{W}_B)), \quad (33)$$

where

$$\begin{aligned} \Omega^{\text{mi}} &\stackrel{\text{def}}{=} \left\{ (\widetilde{W}_A, \widetilde{W}_B) \in \widetilde{C}_A \times \widetilde{C}_B \mid (27) \text{ and } (28) \text{ hold} \right\} \\ &= \left\{ (\widetilde{W}_A, \widetilde{W}_B) \in \widetilde{C}_A \times \widetilde{C}_B \mid \widetilde{W} = \widetilde{W}_A + \widetilde{W}_B \right\}. \end{aligned}$$

**Remark 3.5.** In [16] the authors highlight that the maximization processes (32) and (33) are called sup-convolution operations. This operation is well-known in convex analysis, see for instance [31], where it is denoted  $\square$ . It reads

$$\begin{aligned} S^{\text{immi}}(W) &= S_A \square S_B(W) = \sup_{W_A \in C_A} S_A(W_A) + S_B(W - W_A), \\ S^{\text{mi}}(V_0, \widetilde{W}) &= S_A \square_{V_0} S_B(V_0, \widetilde{W}) = \sup_{\widetilde{W}_A \in \widetilde{C}_A} S_A(V_0, \widetilde{W}_A) + S_B(V_0, \widetilde{W} - \widetilde{W}_A). \end{aligned}$$

Note that in the miscible case, the sup-convolution is partial and acts only on the mass and energy variables, the volume  $V_0$  being fixed. The sup-convolution shares many properties. Among them, it ensures the existence and the concavity of the equilibrium mixture entropies.

**Remark 3.6.** If one assumes that no mass transfer arises between the two phases, then the phasic masses are fixed. In the case of an immiscible mixture without phase transition, the mixture entropy is also concave with respect to  $V, M, E$  and  $M_A$  as soon as the phasic entropies  $S_k$  are sub-linear, see [23]. As miscible mixture without phase transition is concerned, the mixture entropy is also a concave function of  $(V, M, E)$  since it corresponds to a linear transformation of a concave function.

Following [23], one can prove that the extensive entropies  $S^{\text{immi}}, S^{\text{mi}}: C \rightarrow \mathbb{R} \cup \{-\infty\}$  are PH1 concave functions of  $W \in C$ , where  $C = C_A + C_B$  is a non empty convex cone.

When the equilibrium entropies  $S^\lambda$ ,  $\lambda = \text{immi}$  or  $\text{mi}$ , are differentiable with respect to the volume  $V$  and the energy  $E$ , one can define the mixture temperature, pressure and chemical potential at equilibrium as follows:

$$T \stackrel{\text{def}}{=} \left( \frac{\partial S^\lambda}{\partial E} \Big|_{M,V} (W) \right)^{-1}, \quad p \stackrel{\text{def}}{=} T \frac{\partial S^\lambda}{\partial V} \Big|_{M,E} (W), \quad \mu \stackrel{\text{def}}{=} T \frac{\partial S^\lambda}{\partial M} \Big|_{V,E} (W), \quad (34)$$

leading to the extensive fundamental relation

$$TdS^\lambda = dE + pdV - \mu dM. \quad (35)$$

We now turn to the characterization of the equilibrium. In the definition of the mixture entropy, the maximizer may be not unique but the upper value is unique. Besides if the phasic entropies  $S_k$  are upper semi-continuous, then the supremum is actually a maximum.

**Proposition 3.7.** *When the maximum is reached in the interior of the set of constraints, one has:*

- *the maximum is always characterized by the thermal equilibrium*

$$T = T_A = T_B; \quad (36)$$

- *the maximum defining  $S^{\text{immi}}$  in the **immiscible** case is characterized by*

$$p = p_A = p_B, \quad (37)$$

*while, in the **miscible** case with  $S^{\text{mi}}$ , it corresponds to the Dalton's law [30]:*

$$p = p_A + p_B; \quad (38)$$

- *as phase transition occurs in the miscible and miscible mixture, the equilibrium is also characterized by*

$$\mu = \mu_A = \mu_B. \quad (39)$$

We refer to [1, 16, 22] for a proof for the immiscible case and to [15] for both miscible and immiscible cases. Here we focus on the equilibrium reached in the interior of the domain of  $S^\lambda$ . It can also be reached on the boundary of the set of constraints. In that case, one should characterized the maximum by inequalities on the pressures and temperatures. Since the immiscible case is treated in [23] (considering the mass  $M_A$  is fixed), we just give some intents in the miscible case, that is  $V = V_A = V_B = V_0 > 0$  fixed. If the maximum is reached at the point  $(\bar{M}, \bar{E})$ , the concavity of the entropy leads to

$$\nabla_{(M_A, E_A)}(S_A(V_0, M_A, E_A) + S_B(V_0, M - M_A, E - E_A))((M_A, E_A) - (\bar{M}, \bar{E})) \leq 0. \quad (40)$$

Assume that the energy constraints is saturated, then one has

- either  $\bar{E} = E$  and  $\frac{1}{T_A} - \frac{1}{T_B} \geq 0$ ;

- or  $\bar{E} = 0$  and  $\frac{1}{T_A} - \frac{1}{T_B} \leq 0$ .

Conversely, if the mass constraint is saturated, one gets

- either  $\bar{M} = M$  and  $-\frac{\mu_A}{T_A} + \frac{\mu_B}{T_B} \geq 0$ ;
- or  $\bar{M} = 0$  and  $-\frac{\mu_A}{T_A} + \frac{\mu_B}{T_B} \leq 0$ .

### 3.2. Intensive entropies and geometrical interpretation

We now turn to the definition of intensive quantities. The system is now entirely described by its *intensive specific entropy*  $s$  defined by (5) as a function of  $w \stackrel{\text{def}}{=} (\tau, e) \in (\mathbb{R}^+)^2$  or by its *intensive volumic entropy*  $\sigma$  defined by (11) as a function of  $\tilde{w} \stackrel{\text{def}}{=} (\rho, \varepsilon) \in (\mathbb{R}^+)^2$ . We assume that they comply with Assumptions 2 and 3 respectively.

Let us introduce the volume fraction  $\alpha_k$ , the mass fraction  $\varphi_k$  and the energy fraction  $z_k$  of the phase  $k \in \mathbb{K}$  defined respectively by

$$\alpha_k \stackrel{\text{def}}{=} \frac{V_k}{V}, \quad \varphi_k \stackrel{\text{def}}{=} \frac{M_k}{M}, \quad z_k \stackrel{\text{def}}{=} \frac{E_k}{E}, \quad (41)$$

which belong to  $[0, 1]$ .

- Each phase  $k \in \mathbb{K}$  is depicted by its *specific phasic entropy*  $s_k$  defined by

$$s_k(\tau_k, e_k) \stackrel{\text{def}}{=} S_k(\tau_k, 1, e_k). \quad (42)$$

It is a function of the specific volume  $\tau_k$  and a specific internal energy  $e_k$  such that

$$\tau_k = \frac{V_k}{M_k} = \frac{\alpha_k}{\varphi_k} \tau, \quad e_k = \frac{E_k}{M_k} = \frac{z_k}{\varphi_k} e, \quad (43)$$

satisfying the phasic fundamental relation

$$T_k ds_k = de_k + p_k d\tau_k. \quad (44)$$

- Eventually the phase  $k \in \mathbb{K}$  may be depicted by its *volumic phasic entropy*  $\sigma_k$  defined by

$$\sigma_k(\rho_k, \varepsilon_k) \stackrel{\text{def}}{=} S_k(1, \rho_k, \varepsilon_k). \quad (45)$$

It is a function of the density  $\rho_k$  and the volumic internal energy  $\varepsilon_k$  such that

$$\rho_k = \frac{M_k}{V_k} = \frac{\varphi_k}{\alpha_k} \rho = \frac{1}{\tau_k}, \quad \varepsilon_k = \frac{E_k}{V_k} = \frac{z_k}{\alpha_k} \varepsilon = \frac{z_k}{\alpha_k} \rho e = \frac{e_k}{\tau_k}, \quad (46)$$

satisfying the phasic fundamental relation

$$T_k d\sigma_k = d\varepsilon_k - \mu_k d\rho_k. \quad (47)$$

The extensive constraints (27)-(30) admit their intensive counterparts. The mass and energy conservations (27) and (28) give the following constraints on the fractions:

$$1 = \varphi_A + \varphi_B, \quad (48)$$

$$1 = z_A + z_B. \quad (49)$$

- In the case of an **immiscible** mixture, the extensive volume constraint (29) translates into

$$1 = \alpha_A + \alpha_B, \quad (50)$$

and, associated to (43), leads to

$$\tau = \varphi_A \tau_A + \varphi_B \tau_B, \quad (51)$$

$$e = \varphi_A e_A + \varphi_B e_B. \quad (52)$$

Out of equilibrium, the extensive mixture entropy (31) can be written in intensive specific formulation using the homogeneity property of the extensive entropies  $S_k$ ,  $k \in \mathbb{K}$ :

$$s(w_A, w_B) \stackrel{\text{def}}{=} \frac{\mathcal{S}}{M} = \varphi_A s_A(w_A) + \varphi_B s_B(w_B). \quad (53)$$

- In the **miscible** case, the extensive volume constraint (30) translates into

$$1 = \alpha_A = \alpha_B, \quad (54)$$

and, associated to (46), we have

$$\rho = \rho_A + \rho_B, \quad (55)$$

$$\varepsilon = \varepsilon_A + \varepsilon_B. \quad (56)$$

Out of equilibrium, the extensive mixture entropy (31) can be written in intensive volumic formulation using the homogeneity property of the extensive entropies  $S_k$ ,  $k \in \mathbb{K}$ :

$$\sigma(\tilde{w}_A, \tilde{w}_B) \stackrel{\text{def}}{=} \frac{\mathcal{S}}{V} = \sigma_A(\tilde{w}_A) + \sigma_B(\tilde{w}_B). \quad (57)$$

We now turn to the description of the intensive mixture entropy at equilibrium. The optimization processes (32) and (33) can be derived in an intensive formulation as follows, leading to an elegant geometrical interpretation.

- As the **immiscible** case is considered, at equilibrium, the optimization problem (32) reads under its intensive formulation

$$s^{\text{immi}}(\tau, e) = \max_{\substack{1 = \varphi_A + \varphi_B \\ w = \varphi_A w_A + \varphi_B w_B}} (s(w_A, w_B)), \quad (58)$$

under constraints (48), (51) and (52). One observes that the formulation (58) defines nothing else but the concave hull of the maximum of the two partial entropies  $s_k$ ,  $k \in \mathbb{K}$ :

$$s^{\text{immi}}(w) = \text{concave hull of } \max (s_A(w), s_B(w)). \quad (59)$$

Therefore *the mixture entropy is concave but not strictly concave*. In a region corresponding to the vapor-liquid equilibrium, the graph of  $s^{\text{immi}}$  contains a line of plane. This characterization is highlighted for instance in [1, 3, 11, 12, 15, 16].

- In the **miscible** case, the total volume  $V$  is fixed and equal to  $V_0 > 0$ . Then the optimization process allows to define the *volumic* mixture entropy as

$$\sigma^{\text{mi}}(\tilde{w}) = \max_{\tilde{w} = \tilde{w}_A + \tilde{w}_B} (\sigma(\tilde{w}_A, \tilde{w}_B)), \quad (60)$$

under the mass and energy conservation constraints (55) and (56). Observe that the intensive mixture entropy remains a sup-convolution operation of the phasic entropies

$$\sigma^{\text{mi}}(\rho, \varepsilon) = \sigma_A \square \sigma_B(\rho, \varepsilon), \quad (61)$$

see Remark 3.5. The properties of the sup-convolution operation ensure that if the phasic entropies are strictly concave, then *the mixture entropy is strictly concave* as well [18, 31]. Note that the thermodynamic equilibrium is still characterized by Proposition 3.7 that is the equilibrium of the temperature and the Dalton's law on the pressures.

The equilibrium entropies properties are stated in the following statement:

**Proposition 3.8.** *Consider the thermodynamic equilibrium with the equilibrium temperature  $T$  and the equilibrium chemical potential satisfying (36) and (39), i.e.  $T = T_A = T_B$  and  $\mu = \mu_A = \mu_B$ .*

- *If the mixture is **immiscible**, the equilibrium pressure  $p$  satisfies (37) and the intensive equilibrium entropy  $s^{\text{immi}}$  is a concave function of  $(\tau, e)$  and satisfies*

$$T ds^{\text{immi}} = de + pd\tau. \quad (62)$$

- *If the mixture is **miscible**, the equilibrium pressure  $p$  complies with the Dalton's law (38)  $p = p_A + p_B$  and the volumic equilibrium entropy of the mixture is a concave function of  $(\rho, \varepsilon)$  and satisfies*

$$T d\sigma^{\text{mi}} = d\varepsilon - \mu d\rho. \quad (63)$$

*Proof.*

- We multiply the phasic Gibbs relations (44)  $T_k ds_k = de_k + p_k d\tau_k$  by the mass fractions  $\varphi_k$  and we sum over  $k \in \mathbb{K}$ :

$$\begin{aligned} 0 &= \sum_{k \in \mathbb{K}} (\varphi_k T_k ds_k - \varphi_k de_k - \varphi_k p_k d\tau_k) \stackrel{(36)(37)}{=} T \sum_{k \in \mathbb{K}} (\varphi_k ds_k) - \sum_{k \in \mathbb{K}} (\varphi_k de_k) - p \sum_{k \in \mathbb{K}} (\varphi_k d\tau_k) \\ &= T \sum_{k \in \mathbb{K}} d(\varphi_k s_k) - \sum_{k \in \mathbb{K}} d(\varphi_k e_k) - p \sum_{k \in \mathbb{K}} d(\varphi_k \tau_k) - \sum_{k \in \mathbb{K}} (T s_k - e - p \tau_k) d\varphi_k \\ &= T ds^{\text{immi}} - de - pd\tau + \sum_{k \in \mathbb{K}} \mu_k d\varphi_k \stackrel{(39)}{=} T ds^{\text{immi}} - de - pd\tau + \mu d \left( \sum_{k \in \mathbb{K}} \varphi_k \right) \\ &= T ds^{\text{immi}} - de - pd\tau + \mu d(1) = T ds^{\text{immi}} - de - pd\tau. \end{aligned}$$

- We sum the phasic relations (47) for  $k \in \mathbb{K}$ : since the temperatures are equal, see (36), as well as the chemical potentials, see (39), then

$$0 = \sum_{k \in \mathbb{K}} (T_k d\sigma_k - d\varepsilon_k - \mu_k d\rho_k) \stackrel{(36)(39)}{=} T d\sigma^{\text{mi}} - d\varepsilon - \mu d\rho,$$

and

$$\begin{aligned} \sum_{k \in \mathbb{K}} p_k &= \sum_{k \in \mathbb{K}} (\rho_k \mu_k + T_k \sigma_k - \varepsilon_k) \\ &\stackrel{(36)(39)}{=} \mu \left( \sum_{k \in \mathbb{K}} \rho_k \right) + T \left( \sum_{k \in \mathbb{K}} \sigma_k \right) - \left( \sum_{k \in \mathbb{K}} \varepsilon_k \right) \stackrel{(55)(56)(57)}{=} \mu \rho + T \sigma - \varepsilon = p \end{aligned}$$

□

#### 4. HOMOGENEOUS EQUILIBRIUM FOR TWO-PHASE FLOWS

We now take into account the dynamic of the two-phase mixture assuming that the two-phases have the same velocity. The aim of this section is to provide an Homogeneous Equilibrium Euler's system, called HEM model. The HEM model has to be endowed with appropriate closure laws in agreement with the thermodynamic equilibria studied in Section 3. Emphasis is given to the hyperbolicity and the entropy structure of the models. For the sake of simplicity we consider mono-dimensional case.

##### 4.1. Immiscible case

At thermodynamic equilibrium the **immiscible** two-phase flow is depicted by the multicomponent Euler system

$$\begin{cases} \partial_t \rho + \partial_x(\rho u) = 0, \\ \partial_t(\rho u) + \partial_x(\rho u^2 + p^{\text{immi}}) = 0, \\ \partial_t(\rho \mathcal{E}) + \partial_x((\rho \mathcal{E} + p^{\text{immi}})u) = 0, \\ \mathcal{E} = \frac{1}{2}u^2 + e, \\ \forall k \in \mathbb{K} : p_k = p_k(\tau_k, e_k), \tau_k = \rho_k^{-1}, \end{cases} \quad (64)$$

where the flow has a density  $\rho$  (we also define the specific volume  $\tau = 1/\rho$ ), a velocity  $u$ , a pressure  $p^{\text{immi}}$ , and an internal energy  $e$  ( $\mathcal{E}$  being the total energy). The phase  $k \in \mathbb{K}$  is depicted by its mass fraction  $\varphi_k$ , its pressure  $p_k$ , its specific volume  $\tau_k$  and its specific internal energy  $e_k$ , see Section 3.2. All the phases evolve at the same velocity  $u$ .

The multicomponent Euler system admits eight equations and has thirteen unknowns which are

$$\{ \rho, u, \mathcal{E}, p^{\text{immi}}, e, \varphi_A, \varphi_B, \tau_A, \tau_B, e_A, e_B, p_A, p_B \}$$

Thus one has to provide five closure laws. The first two closure laws are given by the constraints (51) and (52). The three remaining closure laws are given by Proposition 3.7, that is

$$\begin{cases} T^{\text{immi}} = T_A = T_B, \\ p^{\text{immi}} = p_A = p_B, \\ \mu = \mu_A = \mu_B, \end{cases} \quad (65)$$

where  $T^{\text{immi}}$  and  $p^{\text{immi}}$  are the thermodynamic temperature and pressure of the two-phase flow and  $p^{\text{immi}} = p^{\text{immi}}(1/\rho, e)$ .

To prove the hyperbolicity of the system, one would like to apply the Godunov-Mock theorem using the function  $\eta : (\tau, u, \mathcal{E}) \mapsto -s^{\text{immi}}(\tau, \mathcal{E} - u^2/2)$  as a Lax entropy of the system, see for instance [6, 24, 29]. Unfortunately, since *the equilibrium entropy  $s^{\text{immi}}$  is not a strictly concave function* of its arguments, this argument does not apply in our case. One has to study the eigenvalues of the system and to prove that the speed of sound of the mixture is real. First notice that, for any state vector  $w \stackrel{\text{def}}{=} (\tau, e)$ , the speed of sound of the system can be written as

$$c^2(w) = -\tau^2 T(w) (p(w), -1) H_{s^{\text{immi}}}(w) \begin{pmatrix} p(w) \\ -1 \end{pmatrix}, \quad (66)$$

where  $H_s^{\text{immi}}(w)$  denotes the Hessian matrix of  $s^{\text{immi}}$ . Because  $s^{\text{immi}}$  is the concave hull of the phasic entropies  $s_k$ ,  $k \in \mathbb{K}$ , it follows that

- for all  $w$  in pure phase regions

$$v^\top H_{s^{\text{immi}}}(w)v < 0, \quad \forall v \neq 0, \quad (67)$$

- for all  $w$  in the VLE region

$$\text{it exists a unique } v(w) \neq 0, \text{ such that } (v(w))^\top H_{\text{simmi}}(w)v(w) = 0. \quad (68)$$

The hyperbolicity results is stated in [1, 11, 12] by proving the following theorem:

**Theorem 4.1.** *For any state vector  $w$  in the VLE region,*

$$v(w) \not\equiv (p(w), -1), \quad (69)$$

so that the system (64) is strictly hyperbolic.

## 4.2. Miscible case

We now turn to the miscible case. At thermodynamic equilibrium the **miscible** two-phase flow is depicted by the multicomponent Euler system

$$\begin{cases} \partial_t \rho + \partial_x(\rho u) = 0, \\ \partial_t(\rho u) + \partial_x(\rho u^2 + p^{\text{mi}}) = 0, \\ \partial_t(\rho \mathcal{E}) + \partial_x((\rho \mathcal{E} + p^{\text{mi}})u) = 0, \\ \mathcal{E} = \frac{1}{2}u^2 + \frac{\varepsilon}{\rho}, \\ \forall k \in \mathbb{K} : p_k = p_k(\rho_k, \varepsilon_k). \end{cases} \quad (70)$$

The system (70) admits six equations and thirteen unknowns which are

$$\{\rho, u, \mathcal{E}, p^{\text{mi}}, \varphi_A, \varphi_B, \tau_A, \tau_B, e_A, e_B, p_A, p_B\}.$$

Hence one has to provide seven closure laws. The first two closure laws are given by the constraints (55) and (56). The characterization of the equilibrium provides the five remaining closure laws:

$$\begin{cases} T^{\text{mi}} = T_A = T_B, \\ p^{\text{mi}} = p_A + p_B, \\ \mu^{\text{mi}} = \mu_A = \mu_B, \end{cases} \quad (71)$$

where  $p^{\text{mi}}$  is the pressure of the two-phase flow and  $p^{\text{mi}} = p^{\text{mi}}(\rho, \varepsilon)$ .

The hyperbolicity of the system (70) relies on its entropy structure. Indeed in Section 3.2 we have shown that *the volumic equilibrium entropy is strictly concave as soon as the phasic entropies are*. In this case the Godunov-Mock theorem applies (see for instance [6, 24, 29]) and we have the following theorem:

**Theorem 4.2.** *The strictly convex function  $\eta(\rho, u, \mathcal{E}) \mapsto -\sigma(\rho, \rho \mathcal{E} - \rho u^2/2)$  is a Lax entropy for the system (70) associated to the entropy flux  $f_\eta(\rho, u, \mathcal{E}) = -u\sigma$ . thus the system (70) is strictly hyperbolic.*

## 5. HOMOGENEOUS RELAXATION TWO-PHASE MODELS

The equilibrium two-phase models presented in the previous section are difficult to use for practical computations. Indeed the determination of the equilibrium mixture pressure law may be tedious, and even worse, it is not analytical when considering standard phasic EoS such as stiffened gas law, see for instance [11, 12, 28]. In [15] the authors propose to use a Fast Legendre Transform algorithm to compute admissible mixture pressure laws (see also Remark 3.5). Again the resulting mixture pressure law is not analytical but solely a tabulated one.

In order to avoid the manipulation of a complex pressure law, it is convenient to approximate the equilibrium Euler system by a Homogeneous Relaxation Model (HRM), following [3, 11–13, 16, 19, 20]. It consists in adding convection equations of additional variables and to modify the pressure to make it depend on these variables. Relaxation source terms complete the new convection equations. The relaxation structure should be consistent with the second principle of Thermodynamics. In particular, the source terms have to comply with the entropy growth criterion and bring the system to thermodynamic equilibrium.

### 5.1. Immiscible case

Following [3, 11, 19], the additional variables we consider are the fractions of mass, volume and energy of the phase  $A$ :  $Y \stackrel{\text{def}}{=} (\varphi_A, \alpha_A, z_A)$ . Note that one may have chosen another set of variables as well. When no transfer occurs between the liquid and the vapor, the fractions should be perfectly convected, that is

$$\partial_t Y + u \partial_x Y = 0. \quad (72)$$

Combining with the conservation of total mass and adding the source terms to the fraction equations gives

$$\begin{cases} \partial_t(\rho Y) + \partial_x(\rho u Y) = Q, \\ \partial_t \rho + \partial_x(\rho u) = 0, \\ \partial_t(\rho u) + \partial_x(\rho u^2 + p) = 0, \\ \partial_t(\rho \mathcal{E}) + \partial_x((\rho \mathcal{E} + p)u) = 0, \end{cases} \quad (73)$$

with the closure law

$$p = p(1/\rho, e, \varphi_A, \alpha_A, z_A). \quad (74)$$

The system (73) admits the concave function

$$s : (\tau, e, \varphi_A, \alpha_A, z_A) \mapsto \varphi_A s_A \left( \frac{\alpha_A}{\varphi_A} \tau, \frac{z_A}{\varphi_A} e \right) + \varphi_B s_B \left( \frac{1 - \alpha_A}{1 - \varphi_A} \tau, \frac{1 - z_A}{1 - \varphi_A} e \right),$$

already defined in (53), as an entropy function. Recall that its concavity is deduced from the concavity of the phasic entropies  $s_k$ ,  $k \in \mathbb{K}$ . Moreover it complies with

$$ds = \sum_{k \in \mathbb{K}} \frac{z_k}{T_k} de + \sum_{k \in \mathbb{K}} \alpha_k \frac{p_k}{T_k} d\tau + \sum_{k \in \mathbb{K}} \partial_{\varphi_k} s d\varphi_k + \sum_{k \in \mathbb{K}} \partial_{\alpha_k} s d\alpha_k + \sum_{k \in \mathbb{K}} \partial_{z_k} s dz_k. \quad (75)$$

We can then identify the mixture pressure (74) and the mixture temperature by

$$\begin{aligned} \frac{p(1/\tau, e, \varphi_A, \alpha_A, z_A)}{T(1/\tau, e, \varphi_A, \alpha_A, z_A)} &= \sum_{k \in \mathbb{K}} \alpha_k \frac{p_k \left( \frac{\alpha_k}{\varphi_k} \tau, \frac{z_k}{\varphi_k} e \right)}{T_k \left( \frac{\alpha_k}{\varphi_k} \tau, \frac{z_k}{\varphi_k} e \right)}, \\ \frac{1}{T(1/\tau, e, \varphi_A, \alpha_A, z_A)} &= \sum_{k \in \mathbb{K}} \frac{z_k}{T_k \left( \frac{\alpha_k}{\varphi_k} \tau, \frac{z_k}{\varphi_k} e \right)}, \end{aligned}$$

under the constraints (48), (49) and (50) on the fractions.

Weak entropy solutions of the system satisfy

$$\partial_t(\rho s) + \partial_x(\rho u s) \geq 0. \quad (76)$$



As relaxation towards the equilibrium is infinitely fast, one recovers the equilibrium fractions which satisfy

$$Y^{\text{immi}}(\tau, e) \stackrel{\text{def}}{=} \operatorname{argmax}_{(\varphi_A, \alpha_A, z_A)} s(\tau, e, \varphi_A, \alpha_A, z_A). \quad (77)$$

As a result the equilibrium pressure is

$$p^{\text{immi}}(\tau, e) = p(\tau, e, Y^{\text{immi}}(\tau, e)). \quad (78)$$

We now turn to the definition of the relaxation source terms involved in the fraction equations in the system (73). Actually there is no consensus on their definition [7, 8]. But they must relax towards the correct thermodynamic equilibrium and comply with the entropy growth criterion (76). Following for instance [3, 12, 13, 16], one may choose  $Q(Y) \stackrel{\text{def}}{=} Y^{\text{immi}}(\tau, e) - Y$ . Thus one has

$$\begin{aligned} \partial_t(\rho s) + \partial_x(\rho u s) &= \nabla_Y s \cdot (\partial_t Y + u \partial_x Y) \\ &= \nabla_Y s \cdot (Y^{\text{immi}}(\tau, e) - Y) \\ &\geq s(\tau, e, Y^{\text{immi}}) - s(\tau, e, \varphi_A, \alpha_A, z_A) \geq 0, \end{aligned}$$

by the concavity of the mixture entropy  $s$ . This choice is then in agreement with the entropy growth criterion.

## 5.2. Miscible case

In the miscible case, the volume fractions are equal to one, see equation (54). Then we consider the additional advection equations on the mass and energy fractions of the phase  $A$ ,  $\tilde{Y} \stackrel{\text{def}}{=} (\varphi_A, z_A)$

$$\partial_t \tilde{Y} + u \partial_x \tilde{Y} = 0. \quad (79)$$

Combining with the conservation of total mass and adding the source terms to the fraction equations gives

$$\begin{cases} \partial_t(\rho \tilde{Y}) + \partial_x(\rho u \tilde{Y}) = \tilde{Q}, \\ \partial_t \rho + \partial_x(\rho u) = 0, \\ \partial_t(\rho u) + \partial_x(\rho u^2 + \tilde{p}) = 0, \\ \partial_t(\rho \mathcal{E}) + \partial_x((\rho \mathcal{E} + \tilde{p})u) = 0, \end{cases} \quad (80)$$

with the closure law

$$\tilde{p} = \tilde{p}(1/\rho, e, \varphi_A, z_A). \quad (81)$$

The system (80) admits the concave function

$$\tilde{s}: (\tau, e, \varphi_A, z_A) \mapsto \varphi_A s_A \left( \frac{\tau}{\varphi_A}, \frac{z_A}{\varphi_A} e \right) + \varphi_B s_B \left( \frac{\tau}{1 - \varphi_A}, \frac{1 - z_A}{1 - \varphi_A} e \right),$$

as an entropy function, which complies with

$$d\tilde{s} = \sum_{k \in \mathbb{K}} \frac{z_k}{T_k} de + \sum_{k \in \mathbb{K}} \frac{p_k}{T_k} d\tau + \sum_{k=A,B} \partial_{\varphi_k} s d\varphi_k + \sum_{k \in \mathbb{K}} \partial_{z_k} s dz_k. \quad (82)$$

We can then identify the mixture pressure (81) and the mixture temperature by

$$\frac{\tilde{p}(1/\tau, e, \varphi_A, z_A)}{\tilde{T}(1/\tau, e, \varphi_A, z_A)} = \sum_{k \in \mathbb{K}} \frac{p_k \left( \frac{\tau}{\varphi_k}, \frac{z_k}{\varphi_k} e \right)}{T_k \left( \frac{\tau}{\varphi_k}, \frac{z_k}{\varphi_k} e \right)},$$

$$\frac{1}{\tilde{T}(1/\tau, e, \varphi_A, z_A)} = \sum_{k \in \mathbb{K}} \frac{z_k}{T_k \left( \frac{\tau}{\varphi_k}, \frac{z_k}{\varphi_k} e \right)},$$

under the constraints (48), (49) and (54) on the fractions.

Weak entropy solutions of the system satisfy

$$\partial_t(\rho \tilde{s}) + \partial_x(\rho u \tilde{s}) \geq 0. \quad (83)$$

As relaxation towards the equilibrium is infinitely fast, one recovers the equilibrium fractions which satisfy

$$\tilde{Y}^{\text{mi}}(\tau, e) \stackrel{\text{def}}{=} \operatorname{argmax}_{(\varphi_A, z_A)} \tilde{s}(\tau, e, \varphi_A, z_A). \quad (84)$$

As a result the equilibrium pressure is

$$\tilde{p}^{\text{mi}}(\tau, e) = \tilde{p}(\tau, e, \tilde{Y}^{\text{mi}}(\tau, e)). \quad (85)$$

Like in the immiscible case, a possible choice for the relaxation source terms on the fractions is  $\tilde{Q}(\tilde{Y}) = \tilde{Y}^{\text{mi}}(\tau, e) - \tilde{Y}$  which guarantees the entropy growth criterion thanks to the concavity of the mixture entropy  $\tilde{s}$ .

## 6. CONCLUSION & PERSPECTIVES

We address in this paper the construction of admissible equations of state for two-phase flows. Considering the mixture as a single fluid, it is possible to construct a complete EoS based on the knowledge of an incomplete EoS. This construction allows for instance to build the van der Waals entropy function starting from the pressure law  $(\tau, T) \mapsto p$ . However this law does not comply with convexity constraints required by standard Thermodynamics. Another construction consists in defining the mixture equilibrium entropy as an optimization problem on the phasic entropies. Depending on the mixture is immiscible or not, the mixture entropy corresponds to the concave hull of the phasic entropies or to their sup-convolution. Both operations guarantee concavity properties and, as a consequence, hyperbolicity of the multicomponent Euler system. However the mixture pressure laws are often difficult to compute and are, at best, tabulated EoS, which are difficult to manipulate and CPU consuming. One comes back to the remaining question 4 of the Introduction: “How to balance accuracy and reasonable computing times?” A solution is to consider rather HRM models. Hence the relaxation source terms have to relax towards the correct thermodynamic equilibrium while complying with the entropy growth criterion.

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## A. LEGENDRE TRANSFORM AND MAXWELL RELATIONS — INTENSIVE DESCRIPTION

The entropy  $s$  (or the energy  $e$ ) are not directly physically measurable, whereas certain of the intensive variables (*e.g.*  $T$  or  $p$ ) are. It is consequently convenient to convert the entropy into functions of various intensive variables. Since  $s$  is concave, we can employ ideas from convex analysis to rewrite various formula in terms of the intensive variables  $T$  and  $p$ . The main tool will be de Legendre transform.

Assume that  $f: \mathbb{R}^n \rightarrow (-\infty, +\infty]$  is a convex, lower semicontinuous (l.s.c.) and proper function  $r \mapsto f(r)$ . The Legendre transform of  $f$  is the function  $f^*: \mathbb{R}^n \rightarrow (-\infty, +\infty]$  defined by  $f^*(q) = \sup_{r \in \mathbb{R}^n} (r \cdot q - f(r))$ . The function  $f^*$  is likewise convex, l.s.c. and proper. Furthermore  $(f^*)^* = f$ . Moreover, if  $f$  is  $\mathcal{C}^2$  and strictly convex then, given  $q$ , there exists a unique  $r$  which maximizes  $(r \cdot q - f(r))$ , namely the unique  $r = r(q)$  for

which  $q = \nabla f(r)$ . In this case we have  $f^*(q) = r(q) \cdot q - f(r(q))$  with  $r = r(q)$  solving  $q = \nabla f(r)$ . Furthermore  $\nabla f^*(q) = r + (q - \nabla f(r))\nabla_q r = r$  and so  $r = \nabla f^*(q)$ .

In classic Thermodynamics it is convenient to define a complete EoS as the function  $(s, \tau) \mapsto e$  instead of  $(e, \tau) \mapsto s$  as in the present paper. Because the temperature positivity (8) holds, we can solve  $s = s(e, \tau)$  and use  $e$  as a  $\mathcal{C}^1$  function of  $(s, \tau)$ , then we have  $de = Tds - p\tau$  and the Hessian matrix of  $e$  is positive definite, that is

$$\det(H_e) \stackrel{\text{def}}{=} \frac{\partial^2 e}{\partial \tau^2} \Big|_s \frac{\partial^2 e}{\partial s^2} \Big|_\tau - \left( \frac{\partial^2 e}{\partial \tau \partial s} \right)^2 > 0, \quad \frac{\partial^2 e}{\partial \tau^2} \Big|_s > 0, \quad \frac{\partial^2 e}{\partial s^2} \Big|_\tau > 0. \quad (86)$$

It implies that the internal energy  $e$  is strictly convex in  $(s, \tau)$ . There are three possible Legendre transforms of  $(s, \tau) \mapsto e$ , according to whether we transform: in the variable  $s$  only, in  $\tau$  only, or in  $(s, \tau)$  together. Because it is customary in Thermodynamics to take the opposite of the mathematical Legendre transform (*i.e.*  $-\sup \{ \dots \} = \inf \{ -\dots \}$ ), the relevant formula are actually the following:

- the Helmholtz free energy  $(T, \tau) \mapsto f$  (sometimes noted  $a$ ) is

$$f(T, \tau) = \inf_s (e(s, \tau) - Ts) = e(s, \tau) - Ts,$$

where  $s = s(T, \tau)$  solves  $T = \frac{\partial e}{\partial s} \Big|_\tau$ ; the Helmholtz free energy  $f$  is strictly concave in  $T$ , strictly convex in  $\tau$ ;

- the enthalpy  $(s, p) \mapsto h$  is

$$h(s, p) = \inf_\tau (e(s, \tau) + p\tau) = e(s, \tau) + p\tau,$$

where  $\tau = \tau(s, p)$  solves  $-p = \frac{\partial e}{\partial \tau} \Big|_s$ ; the enthalpy  $h$  is strictly concave in  $p$ , strictly convex in  $s$ ;

- the Gibbs potential (*a.k.a.* chemical potential or free enthalpy)  $(T, p) \mapsto \mu$  is

$$\mu(T, p) = \inf_{(s, \tau)} (e(s, \tau) + p\tau - Ts) = e(s, \tau) + p\tau - Ts,$$

where  $s = s(T, p)$  and  $\tau = \tau(T, p)$  solve  $T = \frac{\partial e}{\partial s} \Big|_\tau$  and  $-p = \frac{\partial e}{\partial \tau} \Big|_s$ ; the Gibbs free energy  $\mu$  is strictly concave in  $(T, p)$ .

These functions are called ‘‘Thermodynamic potentials’’. They are complete EoS describing the behaviour of a system as a function of the so-called ‘‘natural variables’’. In Table 1 we resume these relations. The Maxwell relations are obtained by exploiting the exactness of differentials (*i.e.* the order of the mixed second partial derivatives does not matter).

Using Legendre’s transform of  $(e, \tau) \mapsto s$ , we can compute three other potentials (and associated Maxwell’s relations, see Table 2) called Massieu-Planck Functions [32]:

- the function  $(\frac{1}{T}, \tau) \mapsto j = \inf_e (s(e, \tau) - \frac{e}{T}) = s(e, \tau) - \frac{e}{T}$  where  $e = e(\frac{1}{T}, \tau)$  solves  $\frac{1}{T} = \frac{\partial s}{\partial e} \Big|_\tau$ ;
- the function  $(e, \frac{p}{T}) \mapsto \star = \inf_\tau (s(e, \tau) - \frac{p}{T}\tau) = s(e, \tau) - \frac{p}{T}\tau$  where  $\tau = \tau(e, \frac{p}{T})$  solves  $\frac{p}{T} = \frac{\partial s}{\partial \tau} \Big|_e$ ;
- the function  $(\frac{1}{T}, \frac{p}{T}) \mapsto r = \inf_{(e, \tau)} (s(e, \tau) - \frac{p}{T}\tau - \frac{e}{T}) = s(e, \tau) - \frac{p}{T}\tau - \frac{e}{T}$  where  $\tau = \tau(\frac{p}{T}, e)$  and  $e = e(\frac{p}{T}, \frac{1}{T})$  solve  $\frac{p}{T} = \frac{\partial s}{\partial \tau} \Big|_e$  and  $\frac{1}{T} = \frac{\partial s}{\partial e} \Big|_\tau$ .

The function  $j$  was first introduced by Massieu and it is called Massieu’s potential. The function  $r$  was introduced by Planck and is called Planck’s potential.

Following similar ideas for  $(\rho, \sigma) \mapsto \varepsilon$  and  $(\rho, \varepsilon) \mapsto \sigma$  we obtain tables 3 and 4. The function  $\clubsuit$  is sometimes called Kramer function [32, p.63].

The other functions are unnamed in our knowledge.

Note that, since  $\varepsilon$  is strictly convex, its conjugate function is precisely  $-p$ , function of  $(T, \mu) = \nabla \varepsilon(\sigma, \rho)$ . Note that, since  $\sigma$  is strictly concave, its conjugate function is precisely  $\frac{p}{T}$ , function of  $(-\frac{\mu}{T}, \frac{1}{T}) = \nabla \sigma(\rho, \varepsilon)$ .

TABLE 1. Thermodynamic potentials based on complete EoS  $(s, \tau) \mapsto e$ 

Potential	Natural variables	Conjugate variables		Differential	Maxwell relations
$e$	$(s, \tau)$	$T = \left. \frac{\partial e}{\partial s} \right _{\tau}$	$p = - \left. \frac{\partial e}{\partial \tau} \right _s$	$de = Tds - pd\tau$	$-\left. \frac{\partial p}{\partial s} \right _{\tau} = \left. \frac{\partial T}{\partial \tau} \right _s$
$f = e - Ts$	$(T, \tau)$	$s = - \left. \frac{\partial f}{\partial T} \right _{\tau}$	$p = - \left. \frac{\partial f}{\partial \tau} \right _T$	$df = -sdT - pd\tau$	$\left. \frac{\partial p}{\partial T} \right _{\tau} = \left. \frac{\partial s}{\partial \tau} \right _T$
$h = e + p\tau$	$(s, p)$	$T = \left. \frac{\partial h}{\partial s} \right _p$	$\tau = \left. \frac{\partial h}{\partial p} \right _s$	$dh = Tds + \tau dp$	$\left. \frac{\partial \tau}{\partial s} \right _p = \left. \frac{\partial T}{\partial p} \right _s$
$\mu = e - Ts + p\tau$	$(T, p)$	$s = - \left. \frac{\partial \mu}{\partial T} \right _p$	$\tau = \left. \frac{\partial \mu}{\partial p} \right _T$	$d\mu = -sdT + \tau dp$	$\left. \frac{\partial \tau}{\partial T} \right _p = - \left. \frac{\partial s}{\partial p} \right _T$

TABLE 2. Thermodynamic potentials based on complete EoS  $(e, \tau) \mapsto s$ .  
Note that  $j = -f/T$  and  $r = -\mu/T$ 

Potential	Natural variables	Conjugate variables		Differential	Maxwell relations
$s$	$(e, \tau)$	$\frac{1}{T} = \left. \frac{\partial s}{\partial e} \right _{\tau}$	$\frac{p}{T} = \left. \frac{\partial s}{\partial \tau} \right _e$	$ds = \frac{1}{T}de + \frac{p}{T}d\tau$	$\left. \frac{\partial \frac{p}{T}}{\partial e} \right _{\tau} = \left. \frac{\partial \frac{1}{T}}{\partial \tau} \right _e$
$j = s - \frac{1}{T}e$	$(\frac{1}{T}, \tau)$	$e = - \left. \frac{\partial j}{\partial \frac{1}{T}} \right _{\tau}$	$\frac{p}{T} = \left. \frac{\partial j}{\partial \tau} \right _{\frac{1}{T}}$	$dj = -e d\frac{1}{T} + \frac{p}{T}d\tau$	$\left. \frac{\partial \frac{p}{T}}{\partial \frac{1}{T}} \right _{\tau} = - \left. \frac{\partial e}{\partial \tau} \right _{\frac{1}{T}}$
$\star = s - \frac{p}{T}\tau$	$(e, \frac{p}{T})$	$\frac{1}{T} = \left. \frac{\partial \star}{\partial e} \right _{\frac{p}{T}}$	$\tau = - \left. \frac{\partial \star}{\partial \frac{p}{T}} \right _e$	$d\star = \frac{1}{T}de - \tau d\frac{p}{T}$	$-\left. \frac{\partial \tau}{\partial e} \right _{\frac{p}{T}} = \left. \frac{\partial \frac{1}{T}}{\partial \frac{p}{T}} \right _e$
$r = s - \frac{1}{T}e - \frac{p}{T}\tau$	$(\frac{1}{T}, \frac{p}{T})$	$e = - \left. \frac{\partial r}{\partial \frac{1}{T}} \right _{\frac{p}{T}}$	$\tau = - \left. \frac{\partial r}{\partial \frac{p}{T}} \right _{\frac{1}{T}}$	$dr = -e d\frac{1}{T} - \tau d\frac{p}{T}$	$\left. \frac{\partial \tau}{\partial \frac{1}{T}} \right _{\frac{p}{T}} = \left. \frac{\partial e}{\partial \frac{p}{T}} \right _{\frac{1}{T}}$

TABLE 3. Thermodynamic potentials based on complete EoS  $(\sigma, \rho) \mapsto \varepsilon$ .Note that  $\diamond = \rho f$  and  $\heartsuit \stackrel{(14)}{=} -p$ 

Potential	Natural variables	Conjugate variables		Differential	Maxwell relations
$\varepsilon$	$(\sigma, \rho)$	$T = \left. \frac{\partial \varepsilon}{\partial \sigma} \right _{\rho}$	$\mu = \left. \frac{\partial \varepsilon}{\partial \rho} \right _{\sigma}$	$d\varepsilon = Td\sigma + \mu d\rho$	$\left. \frac{\partial \mu}{\partial \sigma} \right _{\rho} = \left. \frac{\partial T}{\partial \rho} \right _{\sigma}$
$\diamond = \varepsilon - T\sigma$	$(T, \rho)$	$\sigma = - \left. \frac{\partial \diamond}{\partial T} \right _{\rho}$	$\mu = \left. \frac{\partial \diamond}{\partial \rho} \right _T$	$d\diamond = -\sigma dT + \mu d\rho$	$\left. \frac{\partial \mu}{\partial T} \right _{\rho} = - \left. \frac{\partial \sigma}{\partial \rho} \right _T$
$\nabla = \varepsilon - \mu\rho$	$(\sigma, \mu)$	$T = \left. \frac{\partial \nabla}{\partial \sigma} \right _{\mu}$	$\rho = - \left. \frac{\partial \nabla}{\partial \mu} \right _{\sigma}$	$d\nabla = Td\sigma - \rho d\mu$	$- \left. \frac{\partial \rho}{\partial \sigma} \right _{\mu} = \left. \frac{\partial T}{\partial \mu} \right _{\sigma}$
$\heartsuit = \varepsilon - T\sigma - \mu\rho$	$(T, \mu)$	$\sigma = - \left. \frac{\partial \heartsuit}{\partial T} \right _{\mu}$	$\rho = - \left. \frac{\partial \heartsuit}{\partial \mu} \right _T$	$d\heartsuit = -\sigma dT - \rho d\mu$	$\left. \frac{\partial \rho}{\partial T} \right _{\mu} = \left. \frac{\partial \sigma}{\partial \mu} \right _T$

TABLE 4. Thermodynamic potentials based on complete EoS  $(\varepsilon, \rho) \mapsto \sigma$ .Note that  $\blacklozenge = \rho j$  and  $\heartsuit = -T\clubsuit$ , that is  $\clubsuit = -p/T$ 

Potential	Natural variables	Conjugate variables		Differential	Maxwell relations
$\sigma$	$(\varepsilon, \rho)$	$\frac{1}{T} = \left. \frac{\partial \sigma}{\partial \varepsilon} \right _{\rho}$	$\frac{\mu}{T} = - \left. \frac{\partial \sigma}{\partial \rho} \right _{\varepsilon}$	$d\sigma = \frac{1}{T}d\varepsilon - \frac{\mu}{T}d\rho$	$- \left. \frac{\partial \frac{\mu}{T}}{\partial \varepsilon} \right _{\rho} = \left. \frac{\partial \frac{1}{T}}{\partial \rho} \right _{\varepsilon}$
$\blacklozenge = \sigma - \frac{1}{T}\varepsilon$	$(\frac{1}{T}, \rho)$	$\varepsilon = - \left. \frac{\partial \blacklozenge}{\partial \frac{1}{T}} \right _{\rho}$	$\frac{\mu}{T} = - \left. \frac{\partial \blacklozenge}{\partial \rho} \right _{\frac{1}{T}}$	$d\blacklozenge = -\varepsilon d\frac{1}{T} - \frac{\mu}{T}d\rho$	$\left. \frac{\partial \frac{\mu}{T}}{\partial \frac{1}{T}} \right _{\rho} = \left. \frac{\partial \varepsilon}{\partial \rho} \right _{\frac{1}{T}}$
$\spadesuit = \sigma + \frac{\mu}{T}\rho$	$(\varepsilon, \frac{\mu}{T})$	$\frac{1}{T} = \left. \frac{\partial \spadesuit}{\partial \varepsilon} \right _{\frac{\mu}{T}}$	$\rho = \left. \frac{\partial \spadesuit}{\partial \frac{\mu}{T}} \right _{\varepsilon}$	$d\spadesuit = \frac{1}{T}d\varepsilon + \rho d\frac{\mu}{T}$	$\left. \frac{\partial \rho}{\partial \varepsilon} \right _{\frac{\mu}{T}} = \left. \frac{\partial \frac{1}{T}}{\partial \frac{\mu}{T}} \right _{\varepsilon}$
$\clubsuit = \sigma - \frac{1}{T}\varepsilon + \frac{\mu}{T}\rho$	$(\frac{1}{T}, \frac{\mu}{T})$	$\varepsilon = - \left. \frac{\partial \clubsuit}{\partial \frac{1}{T}} \right _{\frac{\mu}{T}}$	$\rho = \left. \frac{\partial \clubsuit}{\partial \frac{\mu}{T}} \right _{\frac{1}{T}}$	$d\clubsuit = -\varepsilon d\frac{1}{T} + \rho d\frac{\mu}{T}$	$\left. \frac{\partial \rho}{\partial \frac{1}{T}} \right _{\frac{\mu}{T}} = - \left. \frac{\partial \varepsilon}{\partial \frac{\mu}{T}} \right _{\frac{1}{T}}$